

# Effects of Support and Rh Additive on Co-Based Catalysts in the Ethanol Steam Reforming Reaction

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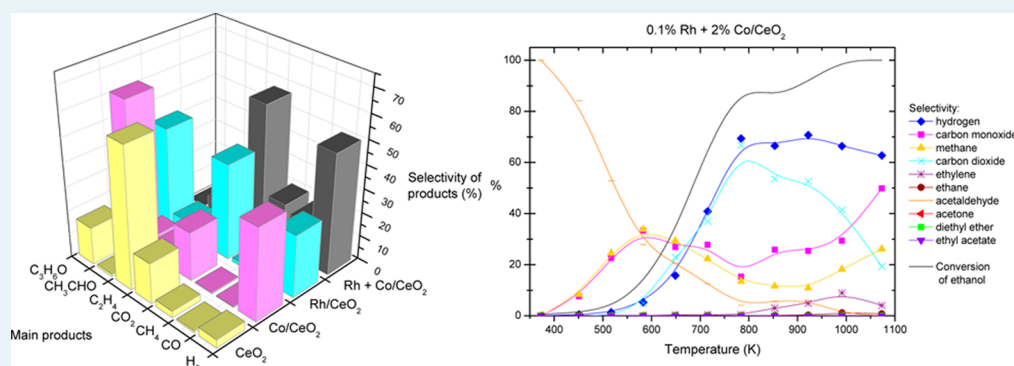
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## Supporting Information



**ABSTRACT:** The effect of the nature of the support and the promotion achieved by a Rh additive on Co-based catalysts in the ethanol steam reforming reaction were studied. The catalysts with 2% Co loading were characterized by temperature-programmed reduction (TPR) and X-ray photoelectron spectroscopy (XPS). In situ diffuse reflectance Fourier-transform infrared spectroscopy (DRIFTS) identified the surface intermediates formed during the reaction, whereas gas phase products were detected by gas chromatography (GC). Upon heating in hydrogen to 773 K, cobalt could not be reduced to  $\text{Co}^0$  on alumina, but on silica the reduction was almost complete. On ceria, half of the Co could be reduced to the metallic state. By the presence of a small amount (0.1%) of Rh promoter, the reduction of both cobalt and ceria was greatly enhanced. For Co on the acidic  $\text{Al}_2\text{O}_3$  support, the dehydration mechanism was dominant, although on the basic  $\text{CeO}_2$  support, a significant amount of hydrogen was also formed. Addition of a small amount of Rh as promoter to the  $\text{Co}/\text{CeO}_2$  catalyst resulted in a significant further increase in the hydrogen selectivity.

**KEYWORDS:** ethanol steam reforming, hydrogen production, cobalt–ceria catalyst, rhodium promoter, ethoxide, DRIFTS, XPS

## 1. INTRODUCTION

Great efforts are currently undertaken to produce hydrogen, for example, for fuel cell applications and for ammonia synthesis by heterogeneously catalyzed processes from renewable sources. This demand inspired studies of the dehydrogenation of oxygenated hydrocarbons.<sup>1–6</sup> In particular, the light alcohol ethanol is an important candidate as a chemical hydrogen carrier. Noble metals, especially Rh, proved to be excellent catalysts for the dehydrogenation reaction,<sup>7,8</sup> but their price is prohibitively high. As an alternative, the less expensive transition metal Co is considered a promising catalyst for the steam reforming of ethanol (SRE).<sup>9–11</sup>

During SRE, acidic supports like  $\text{Al}_2\text{O}_3$  favor dehydration and thereby increase the tendency for coke formation due to the polymerization of ethylene.<sup>12–14</sup> However, on ceria ( $\text{CeO}_2$ ), which is considered to be a basic support, dehydration is

limited and its redox properties hinder coke formation.<sup>6,15</sup> Additionally, ceria promotes the water gas shift (WGS) reaction.<sup>3</sup> Cobalt as catalyst achieves high ethanol conversion and selectivities of over 90% for  $\text{H}_2$  and  $\text{CO}_2$  on  $\text{CeO}_2$  and also on other supports.<sup>16,17</sup> Supported Co catalysts break the C–C bond in adsorbed ethanol.<sup>18</sup> It was found that addition of a  $\text{CeO}_2$  promoter to an unsupported Co powder catalyst stabilizes the more active Co hcp structure and hinders sintering during SRE.<sup>19</sup> Cobalt-based catalysts are also widely used in the reaction of CO and  $\text{H}_2$  to form linear aliphatic hydrocarbons with a broad molecular weight distribution.<sup>20,21</sup> Bartholomew and co-workers pointed out that Fischer–

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Tropsch (FT) synthesis under certain reaction conditions could be structure sensitive.<sup>22</sup> The activity and selectivity appeared to be more closely related to the chemical nature of the support than to the Co dispersion.<sup>23,24</sup> Also, for the reactions of ethanol, similar phenomena should be considered.

Another important physical chemical property of the oxide support is its reducibility. Redox supports such as ceria improve catalyst stability due to their high oxygen storage capacity (OSC) and oxygen mobility. The oxygen exchange capacity of cerium oxide is associated with its ability to reversibly change oxidation states between Ce<sup>4+</sup> and Ce<sup>3+</sup>.<sup>25–27</sup> The easily accessible oxygen can react with carbon species as soon as it forms, and this process keeps the metal surface free of carbon, thus inhibiting deactivation.<sup>9,28,29</sup>

Naturally, the surface properties of the metal and the oxide support and also the metal/oxide interface determine the formation and stability of the intermediates present in the ethanol transformation process. It is generally accepted that the primary step in alcohol activation is the formation of alkoxide.<sup>30</sup> Depending on the particular metal, dehydrogenation and C–C bond scission lead to the formation of alkoxide, oxametallacycle, aldehyde, acyl, and coke on the surface and mostly H<sub>2</sub>, CH<sub>4</sub>, CO, and aldehyde in the gas phase.<sup>30–40</sup> Recent studies suggested that Co<sup>2+</sup> sites are the active centers in SRE, and Co<sup>0</sup> sites are responsible for coke formation;<sup>39,41</sup> however, other authors considered metallic cobalt to play the key role in SRE.<sup>42</sup> High-pressure X-ray photoelectron spectroscopic studies (HPXPS) demonstrated that during the reaction of ethanol with the Co/CeO<sub>2</sub>(111) model catalyst, the amount of Co<sup>2+</sup> decreased drastically with increasing temperature, and at 600 K the majority of Co was metallic; this process was accompanied by a severe reduction of the ceria.<sup>43</sup>

Recently, the effect of adding small quantities of noble metals was investigated on alumina and ceria-zirconia type supports.<sup>44–46</sup> The results showed that the promoting effect of noble metals included a marked decrease in the reduction temperatures of Co<sub>3</sub>O<sub>4</sub> and cobalt oxide surface species due to the hydrogen spillover effect.

In the present work, we aim at finding correlations between the surface properties of supported Co catalysts (such as acid–base character or reducibility with and without a small amount of Rh loading) and their catalytic activity in ethanol steam reforming. The different Co-containing catalysts are characterized by X-ray photoelectron spectroscopy (XPS), X-ray diffraction (XRD), diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS), and temperature-programmed reduction (spectroscopy) (TPR).

## 2. EXPERIMENTAL SECTION

The catalysts were prepared by impregnating the supports (Al<sub>2</sub>O<sub>3</sub> (Degussa P110 C1, 100 m<sup>2</sup>/g), CeO<sub>2</sub> (Alfa Aesar, 43 m<sup>2</sup>/g), and SiO<sub>2</sub> (Cab-O-Sil M5, 200 m<sup>2</sup>/g)) with the aqueous solution of Co(NO<sub>3</sub>)<sub>2</sub> to yield a nominal metal content of 2 wt %, if not mentioned otherwise. The impregnated powders were dried at 383 K, calcined at 973 K, and pressed to pellets. The Rh–Co bimetallic samples were prepared by sequential impregnation (impregnation with Co and calcination first, then the same procedure after impregnation with Rh). Before the measurements, fragments of catalyst pellets were oxidized at 673 K in flowing O<sub>2</sub> for 20 min and reduced at 673 K in flowing H<sub>2</sub> for 60 min in the catalytic reactor.

Catalytic reactions were carried out in a fixed-bed continuous-flow reactor (200 mm long with 8 mm i.d.), which was heated externally. The dead volume of the reactor was filled with quartz beads. The operating temperature was controlled by a thermocouple placed inside

the oven close to the reactor wall, to ensure precise temperature measurement. For catalytic studies, small fragments (about 1 mm) of slightly compressed pellets were used. Typically, the reactor filling contained 50 mg of catalyst. In the reacting gas mixture, the ethanol/water molar ratio was 1:3, if not denoted otherwise. The ethanol–water mixture was introduced into an evaporator with the help of an HPLC pump (Younglin; flow rate: 0.007 mL liquid/min); the evaporator was flushed with Ar flow (60 mL/min). Argon was used as a carrier gas (60 mL/min). The reacting gas mixture-containing Ar flow entered the reactor through an externally heated tube in order to avoid condensation. The space velocity was 60 000 h<sup>−1</sup>.

The analysis of the products and reactants was performed with an Agilent 6890 N gas chromatograph using HP-PLOT Q column. The gases were detected simultaneously by thermal conductivity (TC) and flame ionization (FI) detectors. To increase the sensitivity of CO and CO<sub>2</sub> detection, a methanizer was applied before the detectors.

The rate of ethanol decomposition was defined in terms of conversion,  $X_{\text{ethanol}}$ , whereas the product selectivities were denoted as  $S_{\text{product}}$ . The selectivities,  $S_i$ , toward carbon monoxide, carbon dioxide, methane, ethane, ethylene, acetone, diethyl ether, ethyl acetate, and acetaldehyde were calculated via the carbon balance, defined as the ratio of the product moles to the consumed moles of ethanol, accounting for stoichiometry. The hydrogen selectivity,  $S_{\text{H}_2}$ , was calculated by the hydrogen balance, defined as the molar fraction of hydrogen produced to the total hydrogen in the products:

$$S_i = \frac{x_i n_i}{\sum_j x_j n_j} S_{\text{H}_2} = \frac{2x_{\text{H}_2}}{\sum_j x_j n_j}$$

where  $x_i$  and  $x_{\text{H}_2}$  denote the mole fraction of product ( $i$ ) and H<sub>2</sub>, respectively;  $n_i$  in  $S_i$  is the number of carbon atoms in each molecule of the carbon-containing product ( $j$ ), whereas  $n_j$  in  $S_{\text{H}_2}$  is the number of hydrogen atoms in each molecule of the hydrogen-containing product ( $j$ ). The summation goes for all products.

The amount and the reactivity of surface carbon formed in the catalytic reactions were determined by temperature-programmed hydrogenation. After performing the reactions of ethanol–water mixtures at 823 K for 120 min, the reactor was flushed with Ar at the reaction temperature; then the sample was cooled to 373 K, the Ar flow was changed to H<sub>2</sub>, and the sample was heated up to 1173 K with a 10 K/min heating rate. The formed hydrocarbons were determined and quantified by gas chromatography.

For XPS studies, the powder samples were pressed into pellets with ca. 1 cm diameter and a few tenth of mm thickness, which were placed into the load lock of the spectrometer. Sample treatments were carried out in a high-pressure cell (catalytic chamber) connected to the analysis chamber via a gate valve. They were pretreated in the same way as described above. After the pretreatment, they were cooled to room temperature in flowing nitrogen. Then, the high-pressure cell was evacuated, and the sample was transferred to the analysis chamber in high vacuum (i.e., without contact to air), where the XP spectra were recorded. As the next step, the sample was moved back into the catalytic chamber, where it was treated with the reacting gas mixture at the reaction temperature with the same experimental conditions as used for the catalytic reaction. XP spectra were taken with a SPECS instrument equipped with a PHOIBOS 150 MCD 9 hemispherical electron energy analyzer, using Mg K $\alpha$  radiation ( $h\nu = 1253.6$  eV). The X-ray gun was operated at 210 W (14 kV, 15 mA). The analyzer was operated in the FAT mode, with the pass energy set to 20 eV. The takeoff angle of electrons was 20° with respect to surface normal. Typically five scans were summed to get a single spectrum. For data acquisition and evaluation, both manufacturer's (SpecsLab2) and commercial (CasaXPS, Origin) software were used. A charging of several electron volts was experienced for all samples. The binding energy scale was corrected by fixing the Ce 3d u<sup>'''</sup> peak (see below) to 916.8 eV, the Al 2p peak to 74.7 eV, and the Si 2p peak to 103.4 eV, when using the given supports.

The BET surface and pore volume measurements of the catalysts were carried out by a Quantachrome NOVA 3000e instrument using

N<sub>2</sub> adsorption at liquid nitrogen temperature (Table 1). Apparently, the calcination and the presence of Co only slightly modified the

**Table 1. Surface area ( $S_A$ ), Pore Volume, and Pore Radius of Bare Supports and Co- and Rh-Loaded Catalysts**

catalyst <sup>a</sup>	$S_A$ (m <sup>2</sup> /g)	total pore volume (cm <sup>3</sup> /g) at $p/p_0 = 0.99$	average pore radius (nm)
Al <sub>2</sub> O <sub>3</sub>	110	0.910	16.4
2% Co/Al <sub>2</sub> O <sub>3</sub>	90.8	0.715	15.7
SiO <sub>2</sub>	175	1.620	17.5
2% Co/SiO <sub>2</sub>	168	1.430	17.0
CeO <sub>2</sub>	21.5	0.156	14.5
2% Co/CeO <sub>2</sub>	7.4	0.045	12.3
10% Co/CeO <sub>2</sub>	6.8	0.041	11.7
0.1% Rh/CeO <sub>2</sub>	19.8	0.148	14.2
0.1% Rh+2% Co/CeO <sub>2</sub>	7.6	0.050	13.3

<sup>a</sup>The samples were calcined at 973 K.

surface area of Al<sub>2</sub>O<sub>3</sub>- and SiO<sub>2</sub>-supported catalysts. On the other hand, the calcination led to a significant loss of area for CeO<sub>2</sub>, which was further lowered due to Co loading. The porosity of CeO<sub>2</sub> supported catalysts was very poor compared to alumina and silica supports, in accordance with previous studies.<sup>10</sup> The temperature-programmed reduction (TPR) was carried out in a BELCAT-A apparatus using a reactor (quartz tube with 9 mm outer diameter) that was externally heated. Before the measurements, the catalyst samples were treated in oxygen at 673 K for 30 min. Thereafter, the sample was cooled in flowing Ar to room temperature and equilibrated for 15 min. The oxidized sample was flushed with Ar containing 10% H<sub>2</sub>, the reactor was heated linearly at a rate of 20 K/min up to 1373 K, and the H<sub>2</sub> consumption was detected by a thermal conductivity detector (TCD).

In situ DRIFTS (diffuse reflectance infrared Fourier transform spectroscopy) was used to examine the adsorbed species on the catalysts during the catalytic reactions. The system consists of an FTIR spectrometer (Bio-Rad 135) equipped with a diffuse reflectance attachment (Thermo Scientific) with BaF<sub>2</sub> windows. Following the aforementioned pretreatment steps, the sample was cooled to room temperature under helium flow, and a background file of the pretreated sample was registered.

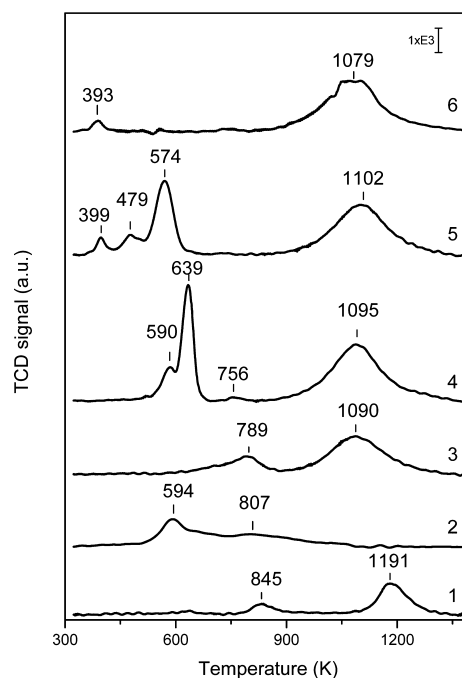
At room temperature, the ethanol steam reforming feed with a water-to-ethanol molar ratio of 3:1 was introduced to the DRIFTS cell. The feed stream was obtained by flowing helium through an evaporator connected to an infusion pump (Econoflow 84 with the flow rate 0.3 mL liquid/h). The tubes were externally heated to avoid condensation. The catalyst was heated under the reaction feed linearly from room temperature to 873 K with a heating rate of 20 K/min, and IR spectra were measured in 50 K intervals. All spectra were recorded between 4000 and 900 cm<sup>-1</sup> at a resolution of 4 cm<sup>-1</sup>. Typically 32 scans were registered; the spectrum of the pretreated catalysts was used as a background. The whole optical path—with the exception of the IR cell—was purged with CO<sub>2</sub>- and H<sub>2</sub>O-free air generated by a Balston purge gas generator. Due to the short optical path within the DRIFTS cell, the contribution of the reactant gases was negligibly small, and from gas phase products only the most intense features were observable.

The XRD study was carried out on a Rigaku Miniflex II powder X-ray diffractometer equipped with a Cu K $\alpha$  radiation source ( $\lambda = 0.15418$  nm) by applying a scanning rate of 4 deg/min in the  $2\theta$  range of 3–80°.

### 3. RESULTS AND DISCUSSION

**3.1. Temperature-Programmed Reduction.** TPR profiles were obtained to investigate the oxide phases present on the catalysts and to evaluate their thermal stability. The reducibility of supported Co catalysts depends on the nature of

oxide support, the calcination temperature, and the amount of metal loading. Figure 1 shows the reduction profiles of the

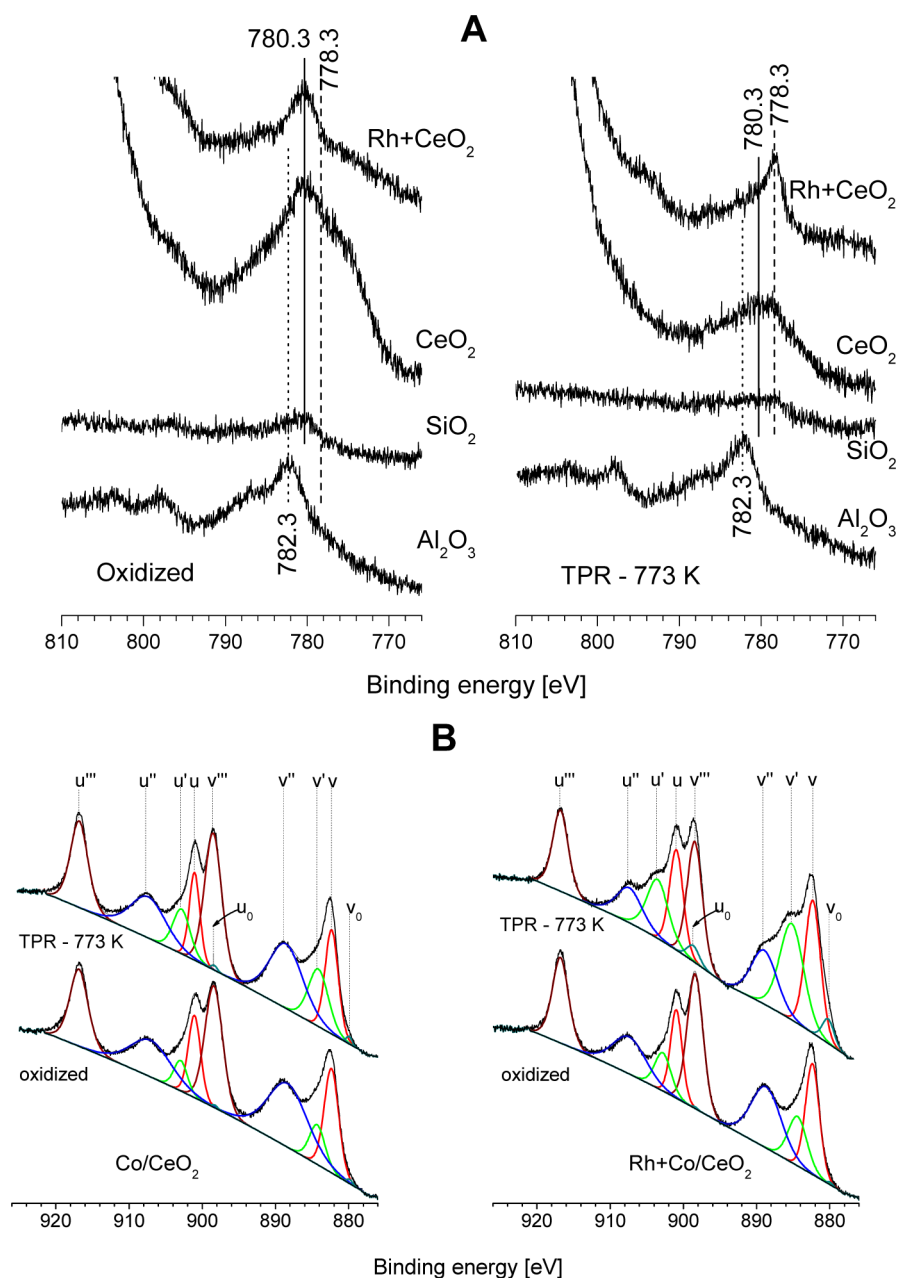


**Figure 1.** TPR profile of 2% Co/Al<sub>2</sub>O<sub>3</sub> (1), 2% Co/SiO<sub>2</sub> (2), CeO<sub>2</sub> (3), 2% Co/CeO<sub>2</sub> (4), 0.1% Rh + 2% Co/CeO<sub>2</sub> (5), and 0.1% Rh/CeO<sub>2</sub> (6). The oxidized sample was flushed with Ar containing 10% H<sub>2</sub>, and the reactor was heated linearly at a rate of 20 K/min up to 1373 K.

catalysts containing 2% Co using different supports. In the case of CeO<sub>2</sub>, for comparison, the TPR curve of the pure oxide and those of the 0.1% Rh + 2% Co/CeO<sub>2</sub> and 0.1% Rh/CeO<sub>2</sub> catalysts are also displayed in Figure 1.

The TPR profile of Co/Al<sub>2</sub>O<sub>3</sub> at 2% Co content (spectrum 1) exhibits a peak at 845 K and a high-temperature reduction peak at 1191 K. For higher metal contents or lower calcination temperatures, the low-temperature reduction peak appears at somewhat lower temperature (~800 K). Some authors suggest that this peak corresponds to the reduction of large crystalline Co<sub>3</sub>O<sub>4</sub> particles to Co<sup>0</sup> via CoO formation,<sup>47</sup> although others ascribe it only to the reduction of Co<sub>3</sub>O<sub>4</sub> to CoO.<sup>48</sup> The higher temperature peak is consistently assigned to the reduction of Co<sup>3+</sup> and Co<sup>2+</sup> species, which are highly dispersed on the surface and strongly interact with alumina,<sup>47,48</sup> forming aluminates during the calcinations process. The latter process is further confirmed by our XRD measurements.

For Co/SiO<sub>2</sub> with Co loadings of 2% (spectrum 2) (also for loadings of 10%, data not shown), an intense reduction peak was observed between 500 and 650 K, with the peak maximum at 600 ± 20 K. In addition, a small, broad peak was detected around 800 ± 100 K. The H<sub>2</sub>-TPR profiles for a 2% Co/SiO<sub>2</sub> sample in Figure 1, with the rate maximum of the low temperature peak at 594 K, agrees well with literature data.<sup>49</sup> This H<sub>2</sub> consumption is assigned to the reduction of Co<sub>3</sub>O<sub>4</sub> to Co<sup>0</sup>.<sup>50,51</sup> The weak high-temperature peak at 807 K has been ascribed to surface cobalt hydroxy silicates, which develop upon reaction of surface silanol groups of the silica support, and is indicative of a limited metal–support interaction.<sup>49,52</sup> For a Co loading of 10%, the peak assigned to reduction of oxide to



**Figure 2.** (A) Co 2p XP spectra on different supports after oxidation at 673 K for 30 min (left) and after linear heating (20 K/min) in H<sub>2</sub> to 773 K (right). (B) Ce 3d XP spectra taken on a 2% Co/CeO<sub>2</sub> catalyst (left) and 0.1% Rh + 2% Co/CeO<sub>2</sub> catalyst (right) after oxidation at 673 K and after reduction with the TPR run up to 773 K.

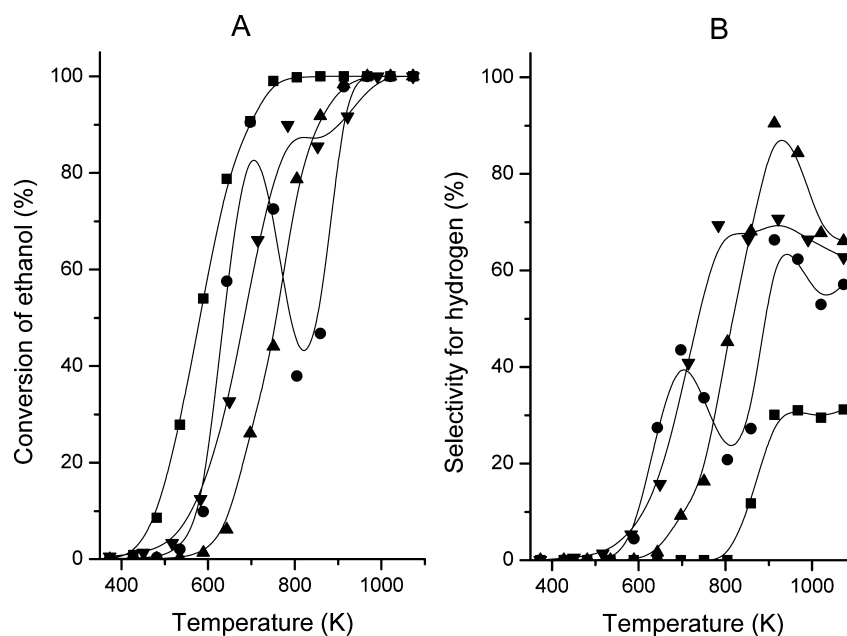
metallic Co increased significantly and shifted to 643 K (not shown).

Before we move on to the loaded support, we first discuss the data for the pure CeO<sub>2</sub> support (spectrum 3). They display an asymmetric low-temperature feature with a rate maximum at 789 K (note that its intensity gradually decreased as TPR runs were repeated, indicating a reduction in the surface region). In addition, a pronounced high-temperature peak was detected at 1090 K, attributed to the bulk reduction of the ceria support.<sup>53</sup>

Upon adding 2% Co on the CeO<sub>2</sub> (spectrum 4), the high-temperature feature shows a slight increase, and an additional well-resolved doublet appeared, with peak maxima at 590 and 639 K. This characteristic reduction profile has been previously reported for Co/CeO<sub>2</sub><sup>10</sup> and is consistent with a stepwise reduction scheme, first from Co<sub>3</sub>O<sub>4</sub> to CoO and thereafter

from CoO to metallic Co. Because XPS results on 2% Co/CeO<sub>2</sub> (see below) indicated that the reduction of Co is not complete up to 773 K, the slightly increased high temperature feature at 1095 K probably also contains a contribution from the reduction of Co. When the Co loading was increased to 10%, the intensity of the low-temperature doublet increased and shifted to lower temperatures (563 and 626 K, not shown), and the peak at around 756 K significantly intensified.

Adding a small amount of noble metal (0.1% Rh) to 2% Co-supported on CeO<sub>2</sub> (spectrum 5) altered the TPR profile significantly. The doublet moved to lower temperatures (479 and 574 K), with the shift being more pronounced for the lower temperature peak (i.e., for reduction of Co<sub>3</sub>O<sub>4</sub> to CoO). The peak observed at 756 K without the noble metal promoter practically disappeared, and the high temperature peak (1102



**Figure 3.** Conversion of ethanol (A) and selectivity of hydrogen (B) as a function of temperature on 2% Co/Al<sub>2</sub>O<sub>3</sub> (■), 2% Co/SiO<sub>2</sub> (▲), 2% Co/CeO<sub>2</sub> (●), and 0.1% Rh + 2% Co/CeO<sub>2</sub> (▼) catalysts.

K) remained more or less unchanged. In addition, a new peak appeared at 399 K, corresponding to the reduction of Rh particles.<sup>53</sup> From these observations, we conclude that small amounts of Rh facilitate the reduction of cobalt oxides. Similar noble metal effects were observed for alumina-supported cobalt Fischer–Tropsch catalysts.<sup>46,54</sup> Presumably, this effect can be attributed to the spillover of hydrogen from Rh to CoO<sub>x</sub>. Similar promoter effects were also found to play an important role on alumina-supported Co/Re, Co/Pt, and Co/Ru catalysts<sup>46</sup> and on other bimetallic cobalt based systems.<sup>6,13,55</sup> The TPR profile of 0.1% Rh/CeO<sub>2</sub> without Co content is also displayed in Figure 1 (spectrum 6). The peak for reduction of Rh particles appeared at 393 K.

**3.2. X-ray Photoelectron Spectroscopy.** Figure 2A shows the X-ray photoelectron spectra in the Co 2p region after oxidation (at 673 K for 30 min) and after hydrogen reduction at 773 K (after TPR experiments) on different supports. In literature, for metallic cobalt (Co<sup>0</sup>), an asymmetric Co 2p<sub>3/2</sub> peak is observed at 778.0–778.5 eV, and Co<sup>2+</sup> is characterized by a Co 2p<sub>3/2</sub> peak at 780–781 eV, with a strong satellite at 786–787 eV. The signature of Co<sup>3+</sup> is a Co 2p<sub>3/2</sub> peak at 780–781 eV with no satellite.<sup>29,39,56</sup>

On the alumina support, the Co 2p<sub>3/2</sub> and 2p<sub>1/2</sub> levels were observed at 782.3 and 798.2 eV, respectively, with both spin–orbit split levels displaying satellites (left panel in Figure 2A). We attribute these spectra for Co/Al<sub>2</sub>O<sub>3</sub> to a Co–aluminate spinel, which was obtained previously in the literature.<sup>57</sup> When the sample was reduced up to 773 K, the spectrum did not change significantly (right panel in Figure 2A). Thus, in accordance with the TPR measurements, no metallic cobalt was detected. When 10% Co was applied, the peak positions remained (not shown), indicating that practically all Co is incorporated in the spinel structure in our case. The main conclusion of our TPR and XPS measurements, therefore, is that Co remained oxidized below ~800 K even in reducing atmosphere.

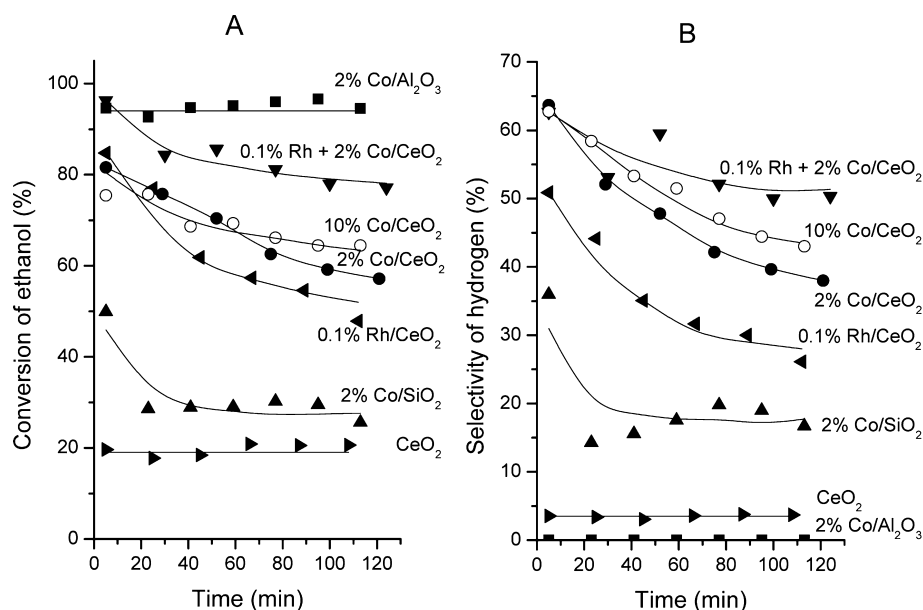
On the SiO<sub>2</sub> support, the intensity of the Co peaks was much smaller than on alumina, which is attributed to the formation of

larger Co/CoO<sub>x</sub> particles. Co appeared in the Co<sup>2+</sup> state after calcination and oxidation, as deduced from the Co 2p<sub>3/2</sub> level at 780.3 eV (Figure 2A, left). When the surface was reduced up to 773 K, in contrast to the alumina support case, the Co 2p<sub>3/2</sub> peak moved to 778.3 eV (Figure 2A, right), which is characteristic for the metallic state.<sup>11,29,39</sup> As the peak was very weak, we cannot exclude that a fraction remained in a partially oxidized state, which may correspond to cobalt hydroxy silicate species.

The 2% Co-supported on ceria showed an intense Co 2p<sub>3/2</sub> peak at 780.3 eV, with a shoulder at lower binding energies, after oxidation (Figure 2A, left). Taking into account the literature data mentioned above, we conclude that the sample contains mainly Co<sup>2+</sup> plus some metallic contribution. When the sample was reduced up to 773 K (Figure 2A, right), the Co 2p<sub>3/2</sub> intensity decreased by more than a factor of 2, and the metallic (778.3 eV) and the Co<sup>2+</sup> components (780.3 eV) exhibited comparable intensities. This indicates that the Co oxide cannot be reduced completely at this temperature. The intensity change after the reduction treatment may reflect a slight encapsulation, sintering or diffusion into the bulk.

The Ce 3d spectra of Co/ceria before and after reduction are shown in Figure 2B. Generally, the Ce 3d region of CeO<sub>2</sub> is rather complex, i.e. it is composed of three doublets, (u<sup>'''</sup>, v<sup>'''</sup>), (u<sup>''</sup>, v<sup>''</sup>) and (u, v), corresponding to the emission from the spin–orbit split 3d<sub>3/2</sub> and 3d<sub>5/2</sub> core levels of Ce<sup>4+</sup>. The three doublets are assigned to different final states: u<sup>'''</sup> (916.8 eV) and v<sup>'''</sup> (898.4 eV) are due to a Ce 3d<sup>9</sup>4f<sup>0</sup> O 2p<sup>6</sup> final state, u<sup>''</sup> (907.7 eV) and v<sup>''</sup> (889.0 eV) to a Ce 3d<sup>9</sup>4f<sup>1</sup> O 2p<sup>5</sup> final state, and u (900.9 eV) and v (882.5 eV) to a Ce 3d<sup>9</sup>4f<sup>2</sup> O 2p<sup>4</sup> final state.<sup>58–61</sup> A minor reduction of Ce<sup>4+</sup> to Ce<sup>3+</sup> is best detectable as the small intensity increase of the u' (903.9 eV) and v' (885.3 eV) peaks and also the weaker u<sub>0</sub> (899.3 eV) and v<sub>0</sub> (880.2 eV) components, which are characteristic of Ce<sup>3+</sup>, after TPR of the 2% Co/CeO<sub>2</sub> catalyst (Figure 2B).

Upon adding a small amount of Rh (0.1%) to the 2% Co/CeO<sub>2</sub> catalyst, pronounced changes of the Co 2p and Ce 3d spectra were detected. After oxidation, a single Co 2p<sub>3/2</sub> feature



**Figure 4.** Conversion of ethanol (A) and selectivity of hydrogen (B) as a function of reaction time at 723 K on 2% Co/Al<sub>2</sub>O<sub>3</sub> (■), 2% Co/SiO<sub>2</sub> (▲), 2% Co/CeO<sub>2</sub> (●), 10% Co/CeO<sub>2</sub> (○), 0.1% Rh + 2% Co/CeO<sub>2</sub> (▼), 0.1% Rh/CeO<sub>2</sub> (◄), and CeO<sub>2</sub> (◄) catalysts.

centered at 780.3 eV appeared due to oxidized cobalt (Figure 2A, left). After TPR up to 773 K (TPR) (or alternatively at 673 K for 1 h), the peak shifted to 778.3 eV (Figure 2A, right), showing a nearly complete reduction of cobalt oxide to metallic cobalt. A similar drastic change was also found in the Ce 3d spectra after reduction (Figure 2B). In particular, the significant intensity increases of the  $v'$  (885.3 eV) and the  $u'$  peaks (903.9 eV) and also the  $u_0$  and  $v_0$  components, clearly indicate a strong reduction of CeO<sub>2</sub>. One of the main messages of this observation thus is that besides the reduction of cobalt on the Rh-doped catalyst, also a considerable reduction of the ceria support occurred during TPR up to 773 K. Presumably, the spillover of H from Rh to CoO<sub>x</sub> and CeO<sub>2</sub> plays an important role.

**3.3. Steam Reforming of Ethanol over Co-Based Catalysts Using Different Supports.** In the catalytic test reaction of ethanol–water steam reforming (1:3 ratio), the conversion of ethanol, the hydrogen selectivity, and the product distribution were studied on cobalt catalysts under the same conditions using different supports and the promoter Rh.

Figure 3A shows the conversion of ethanol during heating (3 K/min) from 373 to 1073 K. The 2% Co/Al<sub>2</sub>O<sub>3</sub> catalyst displays the highest activity, with the reaction starting around 400 K, and almost 100% conversion was reached at 700 K. The 2% Co/CeO<sub>2</sub> catalyst and 0.1% Rh + 2% Co/CeO<sub>2</sub> catalyst were also very active, with the reaction starting at ~500 K. For the former, the increase to full conversion was not monotonous, but after a maximum at ~700 K, there was an intermediate minimum around 800 K. This behavior is assigned either to ethanol desorption through recombinative pathways involving ethoxide or acetaldehyde or to desorption of molecularly adsorbed ethanol, which is present on the 2% Co/CeO<sub>2</sub> surface up to high temperatures (see DRIFT spectra, Figure 8A). On the 2% Co/SiO<sub>2</sub> catalyst, the reaction started around 600 K.

The hydrogen selectivity data, shown in Figure 3B as a function of temperature, showed a value of 70% for the Rh-promoted Co/CeO<sub>2</sub> at ~700 K; on Co/SiO<sub>2</sub>, hydrogen

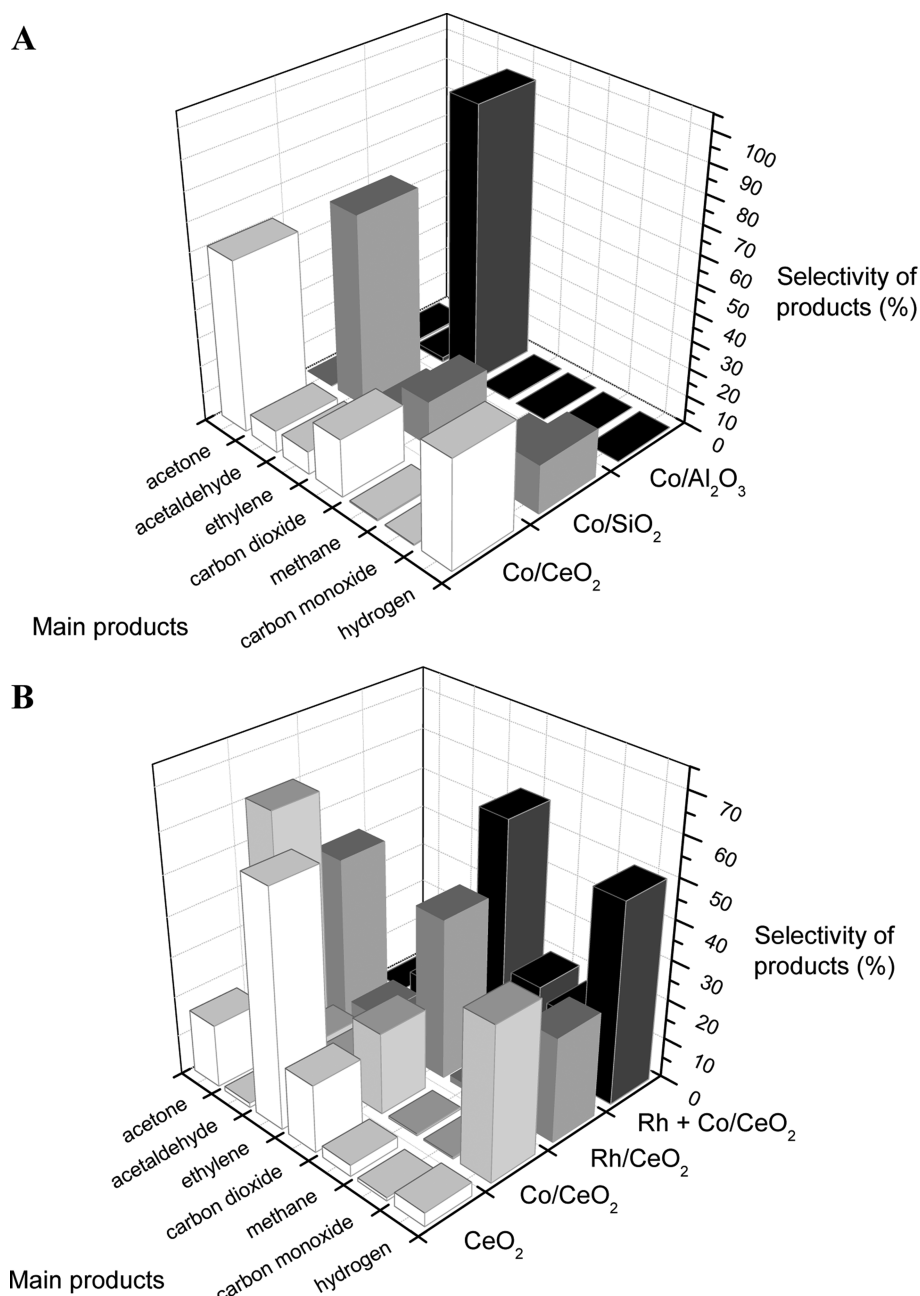
selectivity reached 90%, with an onset at ~650 K. On alumina-supported Co, hydrogen appeared only above 800 K, and the maximum of hydrogen selectivity was only ~30%.

We next address the product distribution (see Supporting Information). On 2% Co/Al<sub>2</sub>O<sub>3</sub> at low temperature (i.e., up to 500 K), when the conversion was low (~5%), the main product was diethyl ether with traces of acetaldehyde and ethylene. The selectivity for ethylene increased between 400 and 600 K, and in the medium temperature range of 650–800 K only ethylene was detected. The lack of O-containing products in this temperature range indicates that, in spite of the presence of a substantial amount of H<sub>2</sub>O in the reaction feed, water also must be a major reaction product, originating from the dehydration reaction, which yields ethylene, too. Above 850 K, ethylene was still the dominant species, but H<sub>2</sub>, CO, CH<sub>4</sub>, acetaldehyde, and a small amount of CO<sub>2</sub> were also formed.

On Co/SiO<sub>2</sub>, below 650 K only acetaldehyde was formed. At higher temperatures, H<sub>2</sub>, CO, CH<sub>4</sub>, CO<sub>2</sub>, and traces of ethylene were also detected, whereas the selectivity for acetaldehyde decreased. The selectivities for H<sub>2</sub>, CO, and CO<sub>2</sub> increased with temperature and exhibited a maximum at 900 K, although the selectivity for CH<sub>4</sub> monotonously increased up to 1100 K. Above 850 K, the main products were hydrogen, CO, CO<sub>2</sub>, and methane.

For the CeO<sub>2</sub> support (without cobalt), initially only acetaldehyde was formed, but between 650 and 850 K, the main product was ethylene. From 700 to 900 K, other products such as acetone, hydrogen, CO<sub>2</sub>, CO, CH<sub>4</sub>, ethane, and ethylene were formed. Above 900 K, hydrogen, CO<sub>2</sub>, ethylene, and smaller amounts of CO and methane were detected.

On 2% Co/CeO<sub>2</sub>, at low conversion up to 500 K, acetaldehyde and acetone were detected. From 500 to 700 K, the acetaldehyde selectivity attenuated the selectivities for H<sub>2</sub> and ethylene, and CO<sub>2</sub> increased moderately; however, the main carbon-containing product still was acetone. Above 700 K, the minimum observed in the conversion and the H<sub>2</sub> selectivity in Figure 3 at ~800 K was mirrored as an increase to a relative maximum for the acetaldehyde selectivity at ~800 K. Above



**Figure 5.** Product distribution in the EtOH + H<sub>2</sub>O (1:3) reaction at 723 K on different oxide-supported Co catalysts (A) and on Co- and Rh-containing ceria-based catalysts (B). Data are plotted at 100 min of reaction time.

800 K, the main products were H<sub>2</sub>, acetaldehyde (only up to 900 K), ethylene, CO<sub>2</sub>, CO, and methane. The main effect of Co as compared to the pure CeO<sub>2</sub> case was manifested in the medium temperature range (650–750 K) as the increased conversion, accompanied by enhanced selectivities for acetone and H<sub>2</sub> at the expense of ethylene.

On the Rh-promoted Co/CeO<sub>2</sub> catalyst, below 550 K the products were acetaldehyde, methane, and CO. Between 600 and 800 K, hydrogen, methane, CO, CO<sub>2</sub>, and acetaldehyde were detected, and above 800 K, hydrogen and CO<sub>2</sub> were dominant, but CH<sub>4</sub> and CO as well as small amounts of ethylene and acetaldehyde were also formed.

For 0.1% Rh/CeO<sub>2</sub> (without Co), initially acetaldehyde and small amounts of CO and methane were formed, but between 650 and 750 K, the main products were hydrogen, acetone, and

CO<sub>2</sub>. Above 850 K, hydrogen and CO<sub>2</sub> were the main products, with CO, methane, and ethylene as minorities. It is worth emphasizing that acetone was not detected at any temperatures on the Rh-promoted Co/CeO<sub>2</sub> catalyst, in spite of the fact that around 700 K it was the major hydrocarbon product adsorbed on 2% Co/CeO<sub>2</sub> and on 0.1% Rh/CeO<sub>2</sub>, and it was easily detectable even on CeO<sub>2</sub> at the same temperature.

To obtain additional information, time-dependent isothermal measurements were carried out at 723 K. Figure 4A displays the ethanol conversion as a function of reaction time, and Figure 4B shows the hydrogen selectivity of the different Co-supported catalysts. The values obtained after 100 min of reaction time largely agree with the conversion and H<sub>2</sub> selectivity data obtained in the TP measurement (Figure 3). At this selected intermediate temperature, the initial ethanol

**Table 2. Vibrational Frequencies for Adsorbed Molecules and Intermediates Formed in Ethanol Steam Reforming on Different Oxide-Supported Co Catalysts**

adsorbed species	vibrational mode	Co/Al <sub>2</sub> O <sub>3</sub>	Co/SiO <sub>2</sub>	CeO <sub>2</sub>	Co/CeO <sub>2</sub>	Rh-Co/CeO <sub>2</sub>	Rh/CeO <sub>2</sub>	refs
		cm <sup>-1</sup>						
ethanol	$\delta(\text{OH})$	1274	(1299)		1269			[14] Ir/Al <sub>2</sub> O <sub>3</sub> [15] Pd/CeO <sub>2</sub> [77] Au/CeO <sub>2</sub>
water	$\delta(\text{H}_2\text{O})$	1641		1640	1621		1650	[9] Co/CeO <sub>2</sub> [35] Pt/Al <sub>2</sub> O <sub>3</sub> [14] Ir/Al <sub>2</sub> O <sub>3</sub>
ethoxide	$\nu_{\text{as}}(\text{CH}_3)$	2976	2979	2966	2975	2975	2973	[9] Co/CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub> [15, 69] CeO <sub>2</sub>
ethoxide	$\nu_{\text{as}}(\text{CH}_2)$	2929	2934	2914	2926			[9] Co/CeO <sub>2</sub> [15, 69] CeO <sub>2</sub>
ethoxide	$\nu_{\text{s}}(\text{CH}_3)$	2898	2898	2878	2861	2883	2886	[15, 69] CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub>
ethoxide	$\delta_{\text{as}}(\text{CH}_3)$	1447	1455			1450		[15] CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub>
ethoxide	$\delta_{\text{s}}(\text{CH}_3)$	1390	1390	1382	1395	1402	1390	[15] CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub>
ethoxide	$\nu(\text{CO})_{\text{mono}}$	1095	1069	1100	1100	1085	1101	[9] Co/CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub>
ethoxide	$\nu(\text{CO})_{\text{bi}}$	1046	1049	1056	1057	1046	1052	[9] Co/CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub>
acetalde- hyde	$\nu(\text{CO})$	1745			1714	1752		[36] Al <sub>2</sub> O <sub>3</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub> [72] Co/CeO <sub>2</sub>
acetate	$\nu_{\text{as}}(\text{OCO})$	1567	1550	1559	1565	1552	1560	[15, 69] CeO <sub>2</sub> [75] Au/CeO <sub>2</sub>
acetate	$\nu_{\text{s}}(\text{OCO})$	1452		1427	1433	1430	1425	[9] Co/CeO <sub>2</sub> [62] Pt/Al <sub>2</sub> O <sub>3</sub> [77] Au/CeO <sub>2</sub>
acetate	$\delta_{\text{s}}(\text{CH}_3)$	1338		1333	1330	1343	1335	[9] Co/CeO <sub>2</sub> [69] CeO <sub>2</sub>
CO	linear					2045		[7] Rh/CeO <sub>2</sub>
CO	bridge						1980	[7] Rh/CeO <sub>2</sub>
croton-aldehyde	$\nu(\text{C}=\text{O})$				1606	1614	1611	[15] Pd/CeO <sub>2</sub> [71] Pt/CeO <sub>2</sub> [77] Au/CeO <sub>2</sub>
	$\nu(\text{C}=\text{C})$							
carbonate	$\nu(\text{OCO})$			1485	1506	1466	1507	[70] CeO <sub>2</sub>
carbonate	$\nu(\text{OCO})$			1389		1338		[70] CeO <sub>2</sub>

conversion was ~80–95% on the alumina and ceria supports, although on Co/SiO<sub>2</sub> it was only ~50%. It is apparent from Figure 4 that the addition of Rh to Co/CeO<sub>2</sub> not only increased the rate for H<sub>2</sub> production but also enhanced the stability of the catalyst. The corresponding product distribution can be seen in Figure 5: panel A shows the products on Al<sub>2</sub>O<sub>3</sub>, SiO<sub>2</sub>, and CeO<sub>2</sub>-supported Co catalysts, and panel B represents the product distribution only on ceria-based catalysts. We plotted average data at 100 min reaction time. The measured values are also listed in the Supporting Information (Table S1). On Co/Al<sub>2</sub>O<sub>3</sub>, only ethylene was detected at this temperature. On Co/SiO<sub>2</sub>, mainly acetaldehyde, hydrogen, carbon monoxide, and carbon dioxide were formed, along with small amounts of methane, ethylene, and traces of ethyl acetate. On Co/CeO<sub>2</sub>, predominantly acetone, H<sub>2</sub>, and CO<sub>2</sub> were formed, plus small amounts of CO, ethylene, and acetaldehyde. For the Rh-modified Co/CeO<sub>2</sub> catalyst, hydrogen and CO<sub>2</sub> are dominant, but CH<sub>4</sub>, CO, and a small amount of acetaldehyde were also formed. The product selectivity was rather stable after

100 min of reaction time. Also for pure CeO<sub>2</sub>, a catalytic activity was observed, yielding C<sub>2</sub>H<sub>4</sub>, CO<sub>2</sub>, acetone, and small amounts of H<sub>2</sub>; the conversion of ethanol was only 25%, whereas it was ~80% on 2% Co/CeO<sub>2</sub>. Adding a small amount of Rh without Co led to a much higher activity compared to pure ceria, and the main products were acetone, CO<sub>2</sub>, and H<sub>2</sub>.

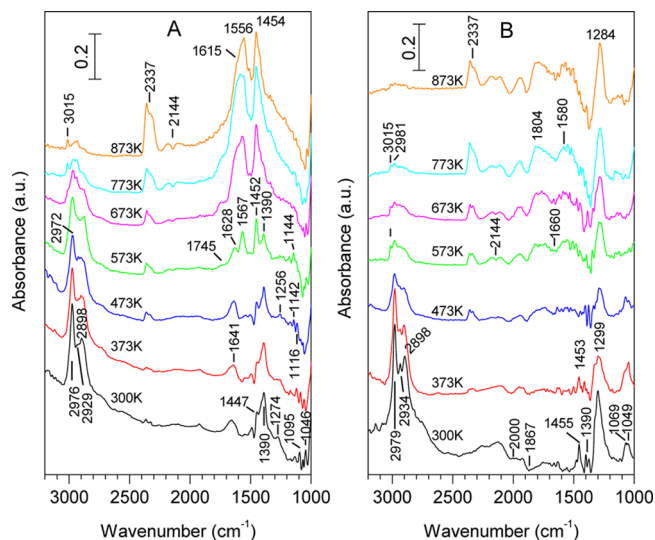
In order to point out the efficiency of the 0.1% Rh promoter, we performed some experiments with a Rh-free 10% Co/CeO<sub>2</sub> catalyst (Figure 4A,B and Table S1). It is clearly shown that the effect of 0.1% Rh is more significant than the increase in Co loading in terms of both conversion and selectivity.

The amount and type of carbon formed in catalytic ethanol steam reforming is an important issue. In agreement with earlier findings,<sup>22</sup> carbon deposits were formed covering both support and cobalt particles, regardless of the type of support used. The extent of coke formation and probably its surface structure depended on the support. After 120 min of reaction at 823 K, the highest quantity, 460  $\mu\text{mol/g}$  of deposited carbon, was determined on Co/SiO<sub>2</sub>, and 344  $\mu\text{mol/g}$  on Co/CeO<sub>2</sub>



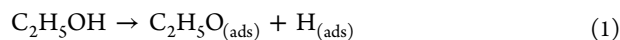
and 55  $\mu\text{mol/g}$  on  $\text{Co}/\text{Al}_2\text{O}_3$  were found. Interestingly, although Rh-promoted  $\text{Co}/\text{CeO}_2$  showed the highest and most stable hydrogen selectivity, the amount of surface carbon, 1135  $\mu\text{mol/g}$ , was higher than that on  $\text{Co}/\text{CeO}_2$ . Previous studies suggested that carbon buildup does not necessarily lead to deactivation.<sup>6,62,63</sup>

**3.4. Infrared Spectra during Ethanol Steam Reforming at Elevated Temperatures.** For catalytic reactions, the exploration of surface species formed during the catalytic processes plays a decisive role in the understanding of the reaction mechanism. Toward this goal, DRIFT spectra were taken at increasing reaction temperatures, in the presence of the reactant mixture/products. The assignment of IR bands and the detailed description is based on the vibrational fingerprints of relevant surface species, which were reported in previous publications.<sup>19,33–36,64–66</sup> The IR bands obtained during the present work and their origin are collected in Table 2. Note that the denoted wave numbers may vary as function of temperature by  $\pm 5\text{ cm}^{-1}$  within one data set. In the following, we will discuss the DRIFT spectra for the different catalysts presented in Figures 6–8; the deduced reactions pathways for all catalysts are summarized in Scheme 1.



**Figure 6.** DRIFT spectra obtained during linear heating (20 K/min) in ethanol–water mixture (1:3) on 2%  $\text{Co}/\text{Al}_2\text{O}_3$  (A) and on 2%  $\text{Co}/\text{SiO}_2$  (B) catalyst.

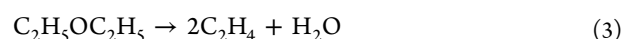
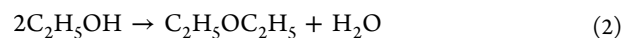
It is generally accepted that in ethanol transformation reactions, the first step is the formation of adsorbed ethoxide ( $\text{C}_2\text{H}_5\text{O}_{(\text{ads})}$ ), according to reaction 1:<sup>6,30,35</sup>



The resulting adsorbed H can form OH groups with lattice O or possibly  $\text{H}_2\text{O}_{(\text{a})}$  with surface OH species.

On  $\text{Co}/\text{Al}_2\text{O}_3$ , the corresponding DRIFTS spectra in Figure 6A show characteristic peaks in the CH stretching frequency range at 2976, 2929, and 2898  $\text{cm}^{-1}$ , which are assigned to the  $\nu_{\text{as}}(\text{CH}_3)$ ,  $\nu_{\text{as}}(\text{CH}_2)$ , and  $\nu_{\text{s}}(\text{CH}_3)$  modes of ethoxide, respectively; the  $\nu_{\text{s}}(\text{CH}_2)$  mode probably overlaps with the  $\nu_{\text{s}}(\text{CH}_3)$  mode 2898  $\text{cm}^{-1}$ . At room temperature, the comparably weak bands of mono- and bidentate ethoxide,  $\nu(\text{CO})_{\text{mono}}$  and  $\nu(\text{CO})_{\text{bi}}$ , can be identified at 1095 and 1046  $\text{cm}^{-1}$  and also the  $\delta_{\text{as}}(\text{CH}_3)$  and  $\delta_{\text{s}}(\text{CH}_3)$  bands at around 1447 and 1390  $\text{cm}^{-1}$ , respectively—see Table 2. The ethoxy

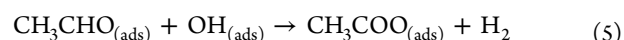
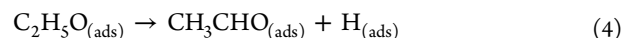
bands decrease with increasing temperature. Our catalytic measurements on  $\text{Co}/\text{Al}_2\text{O}_3$  showed that at 650–800 K the dominant reaction product in the gas phase is ethylene. The Lewis acid–base pairs present on the alumina support are expected to facilitate the dehydration of ethanol (ethoxide) via an  $\text{E}_2$  mechanism.<sup>6,67</sup> Thus, the specific choice of the metal should not alter the reaction paths significantly, and indeed, similar intermediates and products were detected on  $\text{Pt}$ <sup>35,64</sup> or  $\text{Ir}^{14}$  supported on alumina, with the only difference being the absence of adsorbed  $\text{CO}_{(\text{ads})}$  in our study (with bands expected between 1950 and 2120  $\text{cm}^{-1}$ ). The dehydration of ethanol to ethylene presumably occurs at least partly via a two-step dehydration:<sup>6</sup>



The first step (Reaction 2) involves the formation of diethyl ether ( $\text{C}_2\text{H}_5\text{OC}_2\text{H}_5$ ) as a result of intermolecular dehydration from two ethanol molecules. This is followed by a second dehydration of diethyl ether to ethylene (Reaction 3). Indeed, at low temperatures (<550 K) and low conversion diethyl ether is the dominating product in gas phase. At higher temperatures, its conversion to ethylene is very fast, and ethylene is the only product in gas phase. However, we cannot rule out the occurrence of monomolecular dehydration, as it was deduced in DFT calculations.<sup>68</sup>

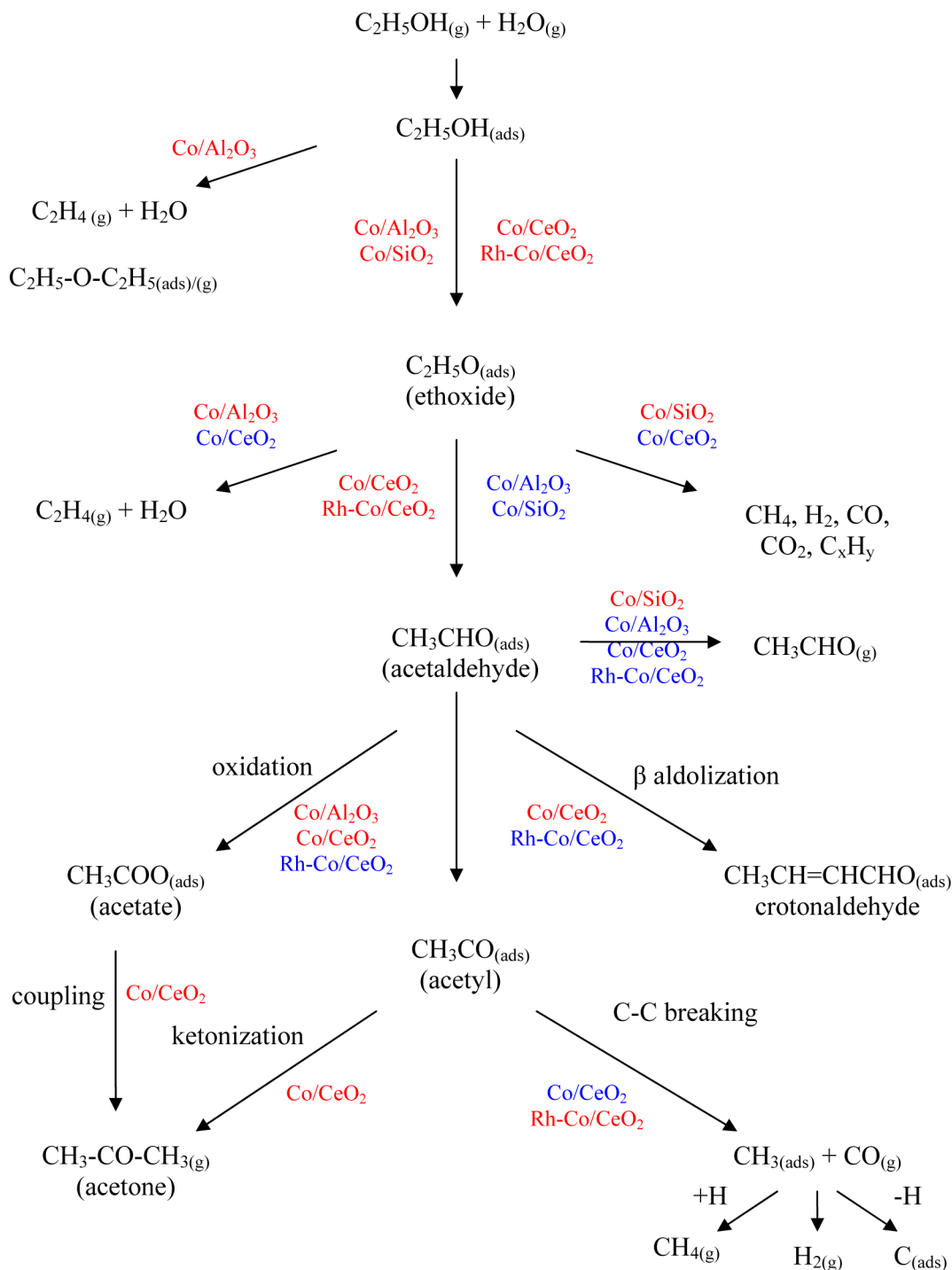
During the reaction of the ethanol–water mixture, chemisorbed water led to the  $\delta(\text{OH})$  band at 1641  $\text{cm}^{-1}$  (Figure 6A), similar to previous observations on  $\text{Co}/\text{CeO}_2$ <sup>9</sup> and on  $\text{Pt}$ <sup>35</sup> or  $\text{Ir}^{14}$  supported by alumina. The 1274  $\text{cm}^{-1}$  band observed at 300 K was attributed to the  $\delta(\text{OH})$  mode of molecularly adsorbed ethanol and was detectable up to 573 K. Molecular ethanol was previously also detected on  $\text{Al}_2\text{O}_3$  and  $\text{Ir}/\text{Al}_2\text{O}_3$  up to 623 K.<sup>14</sup> The assignment of this peak to molecular ethanol is corroborated by the fact that it was very weak, when ethoxide species were produced from  $\text{C}_2\text{H}_5\text{I}$  on  $\text{Al}_2\text{O}_3$ .<sup>66</sup> Molecular ethanol also slightly contributes to the peaks in the CH stretch region and to the low wavenumber peaks of ethoxide.

The adsorbed ethoxy species to a certain extent can further dehydrogenate to acetaldehyde ( $\text{CH}_3\text{CHO}_{(\text{ads})}$ )—Reaction 4. We tentatively assign the weak peak at 1745  $\text{cm}^{-1}$  above 550 K to the  $\nu(\text{CO})$  band in acetaldehyde.<sup>36,65</sup> Furthermore, as shown in the Supporting Information, at low temperatures a small amount of acetaldehyde was detected in gas phase by GC between 350 and 500 K.



Acetaldehyde can be transformed in different ways: it can desorb from the surface above 850 K (see SI), decompose to gas phase methane and CO, or react with hydroxyl groups or with lattice oxygen to form acetate ( $\text{CH}_3\text{COO}_{(\text{ads})}$ )—Reactions 5 and 6).<sup>36,64,65</sup> This stable species is characterized by the  $\nu_{\text{s}}(\text{OCO})$  and  $\nu_{\text{as}}(\text{OCO})$  bands at 1452 and 1567  $\text{cm}^{-1}$  (e.g., at 573 K in Figure 6A), respectively. Its decomposition products at higher temperatures are volatile methane ( $\nu_{\text{as}}(\text{CH})$ , 3015  $\text{cm}^{-1}$ ) and  $\text{CO}_2$  ( $\nu_{\text{as}}(\text{O}=\text{C}=\text{O})$ , 2337  $\text{cm}^{-1}$ ), observed by IR

Scheme 1. Reaction Pathways of SRE on Different Oxide-Supported Co Catalysts (Main Routes Are Colored Red)

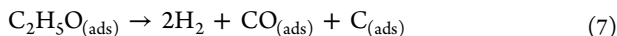


in the gas phase. The proposed reaction pathway is included in Scheme 1.

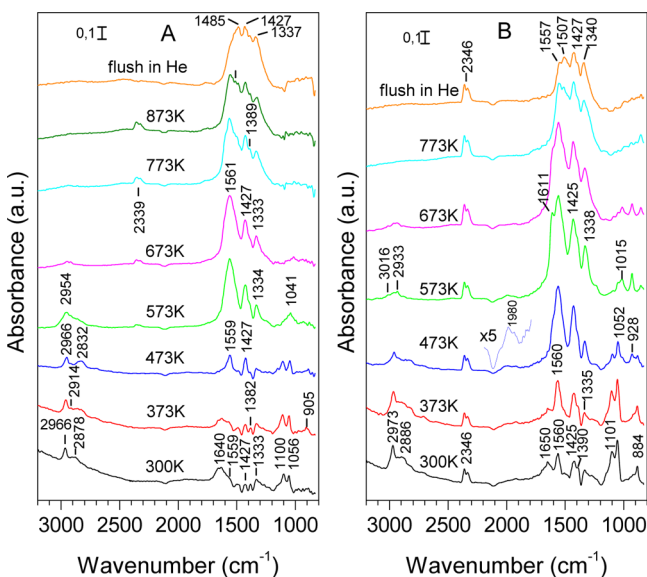
The DRIFT spectra for 2%  $Co/SiO_2$  in the ethanol–water mixture are shown in Figure 6B. The bands at 1069 and 1049  $cm^{-1}$  were assigned to the  $\nu(CO)$  vibrations of mono- and bidentate ethoxide; the peaks at 1455 and 1390  $cm^{-1}$  are attributed to the ethoxide  $\delta_{as}(CH_3)$  and  $\delta_s(CH_3)$  modes, and the peaks at 2979, 2934, and 2898  $cm^{-1}$  to the ethoxide  $\nu_{as}(CH_3)$ ,  $\nu_{as}(CH_2)$ , and  $\nu_s(CH_3)$  modes, respectively. In addition, molecular ethanol may also contribute to the above-mentioned bands. The evaluation of the low wavenumber region of  $Co/SiO_2$  is particularly difficult, because  $SiO_2$  itself

has very strong absorptions at  $\sim 2000$ ,  $\sim 1870$ , and  $\sim 1640$   $cm^{-1}$ , and a sharp absorption edge at  $\sim 1300$   $cm^{-1}$ . Although these features should be accounted for by the background spectrum, they might also change as a function of temperature and thus disturb our spectra. We tentatively assign the features observed at 2000, 1867, 1660  $cm^{-1}$ , and also the intense peak at 1299  $cm^{-1}$  to this effect. The alternative assignment of the 1299  $cm^{-1}$  peak to molecular ethanol (as it was done for  $Co/Al_2O_3$ ) would imply its presence in large amounts on the surface up to 873 K, which is highly unlikely. Upon heating, the ethoxide bands decreased in intensity and disappeared around 573 K. According to our kinetics measurements (Figure 3A), practi-

cally no conversion occurred up to 600 K; only a small amount of acetaldehyde was detected. Above 650 K, besides the continuously decreasing acetaldehyde, hydrogen and CO are the main products in the gas phase. Because SiO<sub>2</sub> is rather inert, surface transformations probably mostly occur on metallic/oxidized Co. Beside acetaldehyde formation (Reaction 4), the primary pathway on Co<sup>0</sup> sites is decarbonylation (Reaction 7), in agreement with the observation on pure metallic cobalt:<sup>41</sup>



Before discussing the results obtained during heating the Co/CeO<sub>2</sub> system in ethanol–water mixture, we briefly summarize the most important findings on clean ceria. Even though the interaction of ethanol and ethanol–water mixtures with CeO<sub>2</sub> was previously studied by other authors<sup>15,69</sup> using IR spectroscopy, we repeated some measurements to have direct comparison with data obtained in the same apparatus. In these experiments, the ethoxide bands were found at 2966, 2914, 2878, 1382, 1100, 1056, and 905 cm<sup>-1</sup>, up to 573 K (Figure 7A,

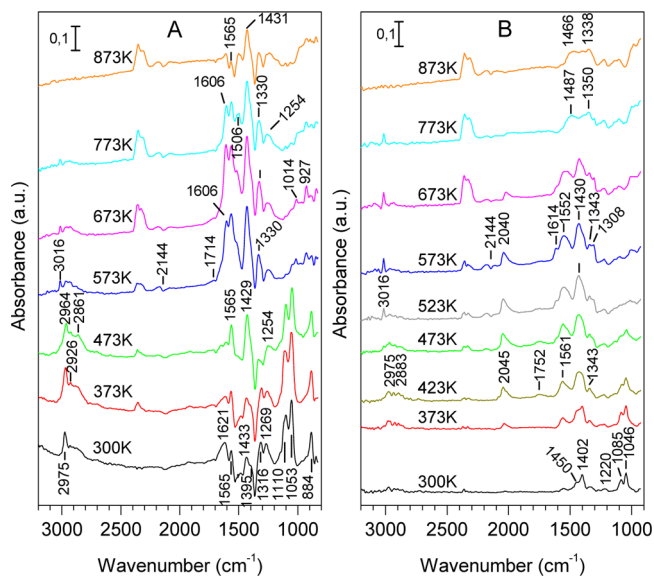


**Figure 7.** DRIFT spectra obtained during linear heating (20 K/min) in ethanol–water mixture (1:3) on CeO<sub>2</sub> (A) and on 0.1% Rh/CeO<sub>2</sub> catalyst (B). The insert of the CO regime at 473 K is magnified by 5.

Table 2). The formation of an acetaldehyde surface species was not observed. These observations are in agreement with recent photoelectron spectroscopic results obtained for ethanol adsorption at low<sup>38</sup> or higher pressures<sup>43</sup> on well-ordered CeO<sub>2</sub>(111) films. The peak at 1640 cm<sup>-1</sup> is assigned to H<sub>2</sub>O<sub>(a)</sub>, which is stable up to 373 K. Upon heating, starting at 473 K, the peaks of acetate species dominated the spectrum (1559, 1427, and 1334 cm<sup>-1</sup>). Please note that minor amounts of this species are already formed at 300 K.<sup>20,69</sup> Starting from 773 K, the acetate groups transformed at least partly into carbonate, as indicated by the peaks at 1485 and 1389 cm<sup>-1</sup>, wavenumbers typical for carbonates.<sup>70</sup> Because the formation of carbonates from acetate is accompanied by slight shifts of the acetate bands,<sup>15,71</sup> it is difficult to determine if any acetate remains on the surface after this process. Nevertheless, the (almost) complete lack of C–H stretching features at T ≥ 773 K suggests that the amount of acetate is rather small (note that for Co/Al<sub>2</sub>O<sub>3</sub> the high thermal stability of acetate groups was

accompanied by easily detectable C–H stretch modes up to 873 K (Figure 6A).

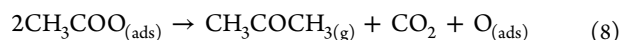
The DRIFTS measurements of the 2% Co/CeO<sub>2</sub> catalyst in the ethanol–water mixture from 300 to 873 K are shown in Figure 8A. The assignment of the peaks was based on previous

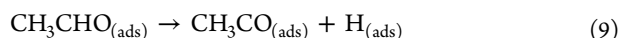


**Figure 8.** DRIFT spectra obtained during linear heating (20 K/min) in ethanol–water mixture (1:3) on 2% Co/CeO<sub>2</sub> (A) and on 0.1% Rh + 2% Co/CeO<sub>2</sub> catalyst (B).

works.<sup>6,9–11,15,28,72,73</sup> The formation of ethoxide (2975, 2926, 2861, 1395, 1100, 1053, and 884 cm<sup>-1</sup>) was observed after the introduction of the reaction feed at room temperature. The band at 1621 cm<sup>-1</sup>, detectable up to 473 K, is assigned to molecular water on the surface, whereas the band at 1269 cm<sup>-1</sup> is attributed to molecular ethanol. In addition, the formation of acetate was detected at this temperature (1565, 1433, 1316 cm<sup>-1</sup>). Although below 573 K small amounts of acetaldehyde were detected in gas phase (Supporting Information), no bands due to adsorbed acetaldehyde were observed (1714 cm<sup>-1</sup>) in this temperature regime. The intensity of ethoxide bands decreased significantly above 473 K. The acetate peaks (1565, 1429, 1330 cm<sup>-1</sup>) increased considerably up to 573 K. It is remarkable that adsorbed CO was not identified in the whole temperature range.

From ~500 K on, acetone was observed in gas phase. The vibrational frequencies of (adsorbed) acetone and acetaldehyde are very close, which renders the discrimination between the two molecules very demanding. Nevertheless, because the peak at 1714 cm<sup>-1</sup> is very small, we conclude that the surface concentration of acetone was very low. From the gas phase product distribution at 723 K in Figure 5, we may conclude that on a basic support such as ceria, ethoxide is dehydrogenated yielding acetaldehyde, which is immediately oxidized to surface acetate species by lattice oxygen or by OH groups (Reactions 5 and 6). The other reaction path is the formation of acetone (CH<sub>3</sub>COCH<sub>3(g)</sub>), which is the dominant product at medium temperatures in our case. According to literature data, acetone can be produced through aldol condensation of acetate (Reaction 8) or via the reaction of acetyl groups (CH<sub>3</sub>CO) with methyl species (Reactions 9–11):<sup>6,72</sup>



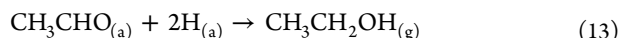


The high acetone production between 500 and 750 K suggests a propensity of the oxide phases for aldol condensation-type reactions because our catalyst contains a significant number of  $\text{Co}^{2+}$  sites, and ceria is still oxidized in this temperature range. We note that product distribution on  $\text{Co}/\text{CeO}_2$  depends on the preparation method and on the support particle size.<sup>11,73,74</sup> In the reaction mechanism, mentioned above, the acetyl intermediate may play an important role, in spite of the fact that it could not be detected by IR during the reaction due to its limited lifetime on the surface. Acetyl species were identified after adsorption of aldehyde on  $\text{Co}/\text{CeO}_2$  at  $1684\text{ cm}^{-1}$ ; this small acetyl band was only present when the aldehyde-covered surface was heated to 335 K.<sup>72</sup> The proposed pathways on  $\text{CeO}_2$ -supported Co catalysts are included in Scheme 1.

From 573 K on, a new feature showed up at  $1606\text{ cm}^{-1}$ , and it was present up to 773 K. We assign this band to  $\nu(\text{C}=\text{O})$  and  $\nu(\text{C}=\text{C})$  modes<sup>73</sup> of adsorbed crotonaldehyde ( $\text{CH}_3\text{CH}=\text{CHCHO}$ ; Reaction 12). This band was also detected and attributed to crotonaldehyde formation on ceria-supported  $\text{Pt}^{14,35}$  and  $\text{Pd}^{15}$  during ethanol steam reforming reaction. The other vibrational modes of adsorbed crotonaldehyde are typically much weaker and are thus not detected.<sup>75</sup> Such kind of four C–C bond formation reactions can be described by  $\beta$ -aldolization of acetaldehyde to crotonaldehyde, as it was shown for the adsorption of acetaldehyde on ceria<sup>36,72</sup> and on ceria-supported Co or noble metals.<sup>7,72</sup>



The formation of gas phase methane ( $3016\text{ cm}^{-1}$ ),  $\text{CO}_2$  ( $2359\text{ cm}^{-1}$ ), and CO ( $2144\text{ cm}^{-1}$ ) was also detected by IR. Above 700 K, the acetate and crotonaldehyde species decomposed, and carbonate species were detected. The band at  $1506\text{ cm}^{-1}$  can certainly be assigned to  $\nu(\text{OCO})$  mode of carbonate (Figure 8A). Because different types of carbonate can coexist on the  $\text{CeO}_2$  surface with peaks partially overlapping with acetate,<sup>70</sup> a clear distinction between these forms is ambiguous. Nevertheless, the weak intensity observed in the  $\nu(\text{C}–\text{H})$  stretching region suggests that the surface concentration of acetate at  $T \geq 773\text{ K}$  is small. The analysis of the gas phase showed methane,  $\text{CO}_2$ , CO, ethylene, acetaldehyde, and  $\text{H}_2$ . At  $\sim 800\text{ K}$ , the conversion and the  $\text{H}_2$  selectivity transiently dropped, which was also seen as an increase in the acetaldehyde selectivity (Supporting Information). A possible reason is the reduced rate of the formation of acetate and acetyl from acetaldehyde, leading to acetaldehyde desorption and recombination to ethanol (Reaction 13)

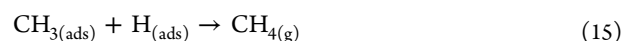
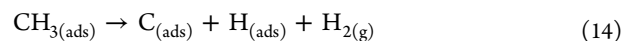


Adding 0.1% Rh to the 2%  $\text{Co}/\text{CeO}_2$  catalyst changed the stability and intensity of intermediates formed during the catalytic reaction, as is evident from the DRIFTS spectra in Figure 8B, which were measured in the presence of the ethanol–water reaction mixture. As established by XPS (see above), Co was transformed into a fully metallic state, and also a significant amount of  $\text{Ce}^{4+}$  was reduced to  $\text{Ce}^{3+}$  after hydrogen reduction up to 773 K (Figure 2). Significantly less

ethoxide was detected, and it was not identified above 523 K. However, adsorbed CO was established already at 373 K: the band at  $2045\text{ cm}^{-1}$  is attributed to CO linearly bonded to Rh or metallic Co.<sup>7,76</sup> From 523 K on, bands for gas phase methane ( $\nu(\text{CH}_3)$  at  $3016\text{ cm}^{-1}$  and  $\delta(\text{CH}_3)$  at  $1308\text{ cm}^{-1}$ ) appeared, and gas phase  $\text{CO}_2$  ( $2359\text{ cm}^{-1}$ ) was detected. Acetaldehyde was observed at  $1752\text{ cm}^{-1}$  on this surface between 423 and 573 K, in line with observations on  $\text{CeO}_2$ -based supports with other noble metals, such as Rh,<sup>7</sup> Pt,<sup>71</sup> Pd,<sup>15</sup> and Au.<sup>77</sup> Between 300 and 673 K, the bands characteristic of acetate bands are detected, albeit with lower intensity than measured on  $\text{Co}/\text{CeO}_2$ . This observation is in agreement with recent findings that Rh promotes the demethanation of acetate to carbonate.<sup>76</sup> In contrast to  $\text{Co}/\text{CeO}_2$ , a weak band from crotonaldehyde at  $1614\text{ cm}^{-1}$  only transiently showed up (523–573 K).

For comparison, DRIFTS measurements were also carried out on the 0.1%  $\text{Rh}/\text{CeO}_2$  catalyst, that is, without cobalt (Figure 7B). The bands characteristic of ethoxide ( $2973$ ,  $2886$ ,  $1390$ ,  $1101$ , and  $1052\text{ cm}^{-1}$ ) are clearly visible at 300 K and decreased above 373 K to finally disappear around 673 K. At room temperature, transiently adsorbed water is also detected ( $1650\text{ cm}^{-1}$ ). Acetate species ( $1560$ ,  $1425$ , and  $1335\text{ cm}^{-1}$ ) were already present at 300 K and grew considerably upon annealing. A (partial) conversion of acetate to carbonate set in at 773 K, as was observed also on  $\text{CeO}_2$ . Crotonaldehyde ( $1611\text{ cm}^{-1}$ ) was detected between 473 and 673 K. In contrast to 0.1% Rh + 2%  $\text{Co}/\text{CeO}_2$ , DRIFTS showed no adsorbed acetaldehyde. It is remarkable that on 0.1%  $\text{Rh}/\text{CeO}_2$ , no linearly adsorbed CO was found, but only a weak band at  $1980\text{ cm}^{-1}$ , attributable to bridge-bonded CO.

From our observations, we conclude that the promoter Rh has at least two different roles in this catalytic system. First, the reduction of Co (and  $\text{CeO}_2$ ) in  $\text{H}_2$  was much more efficient in the presence of Rh (Figures 1 and 2). Because  $\text{H}_2$  is also present as a product, Rh may also help to keep the cobalt in the metallic state. On the other hand, Rh promotes the C–C bond scission reaction of ethanol, producing adsorbed  $\text{CH}_3$ .<sup>7,79</sup> This is in agreement with the fact that on our Rh-containing samples, the selectivities for methane are higher than for the  $\text{CeO}_2$  and  $\text{Co}/\text{CeO}_2$  systems. The reaction of ethanol on metallic cobalt was investigated earlier on a Co foil.<sup>41</sup> In this study, the primary reaction of ethoxide species on the metallic cobalt surface was decarbonylation, presumably through acetaldehyde and acetyl species, producing  $\text{H}_2$ , CO, and carbon. Taking into account these observations, we propose that Reactions 9 and 10 are main reaction steps on 0.1% Rh + 2%  $\text{Co}/\text{CeO}_2$  catalyst, followed by Reactions 14–16:



The products of these reaction steps were detected in gas phase and identified in FTIR spectra. The hydrogen selectivity significantly increased in the presence of a small amount of Rh. The fact that the bimetallic catalyst was the most active and selective for hydrogen production and at the same time it contained the largest fraction of Co in metallic state indicates that metallic cobalt sites are active in the SRE reaction. This conclusion is different from previous suggestions that claim that metallic cobalt sites are mainly active in the carbon formation reaction.<sup>39,41</sup> Most likely, both  $\text{Co}^{2+}$  and metallic Co play roles

in different steps of the SRE reaction. For example,  $\text{Co}^{2+}$  is active in the dehydrogenation of ethanol, whereas metallic sites are particularly active in C–C bond rupture and decarbonylation.

It is worth mentioning that acetone was not detected in the gas phase on Rh-promoted Co/ceria. Acetone was produced through aldol condensation of acetaldehyde; however, this type of reaction requires  $\text{Co}^{x+}$  centers on the surface that are not found in the presence of Rh. Alternatively, in the formation of acetone on  $\text{CeO}_2$ , stoichiometric  $\text{CeO}_2$  is involved, although on Rh–Co/ $\text{CeO}_2$ , cerium ions are also greatly reduced.

The amount of acetate was less on the bimetallic catalyst, indicating that its formation from aldehyde requires more mobile lattice oxygen or OH groups, which are better available without rhodium. As it was shown that Rh catalyzes the demethanation of acetate to form carbonate species and  $\text{CO}_2$ , this process might be an additional hydrogen source (Figure 8B).

#### 4. CONCLUSIONS

The steam reforming of ethanol was studied under identical conditions on different oxide-supported Co-based catalysts with a Co content of 2%. The catalysts were characterized by TPR and XPS, and the conversion of ethanol and the product distribution were determined by gas chromatography between 300 and 1073 K. The transiently formed surface intermediates were identified by in situ DRIFTS. In the following the most important results and conclusions are summarized: (1) Upon TPR to 773 K, cobalt was not reduced to  $\text{Co}^0$  on the alumina support. In contrast, on silica, the majority of cobalt was reduced to the metallic state upon TPR to 773 K. The behavior on ceria is intermediate, with about half of the cobalt was reduced to  $\text{Co}^0$  and half remaining in the  $\text{Co}^{2+}$  state. When a small amount (0.1%) Rh was added to the Co/ceria system, the reducibility dramatically changed, that is, cobalt was completely reduced to the metallic  $\text{Co}^0$  state. In addition, also the ceria support was reduced significantly. The effects can be explained by hydrogen spillover phenomena. (2) On the acidic Co/ $\text{Al}_2\text{O}_3$  catalyst, the highest ethanol conversion was detected at lower temperature, and ethylene was the main product; the hydrogen selectivity was almost zero at 723 K. On Co/ $\text{SiO}_2$  hydrogen, acetaldehyde,  $\text{CO}_2$  and CO were the dominant products, but the activity was rather poor. The basic Co/ $\text{CeO}_2$  catalyst displayed high activity in ethanol steam reforming; at 723 K hydrogen, acetone,  $\text{CO}_2$ , and CO were the main products. In the presence of Rh promoter, the product distribution significantly altered, and the catalyst was more stable; the hydrogen selectivity significantly increased, and  $\text{CO}_2$ , CO,  $\text{CH}_4$ , and some aldehyde were produced, but no acetone formation was detected. (3) As is generally accepted, the first step in ethanol activation is the formation of ethoxide. On Co/ $\text{Al}_2\text{O}_3$ , acetate was identified as the main other adsorbed species in DRIFTS, and no adsorbed CO was detected. On the Co/ $\text{CeO}_2$  catalyst ethoxide, adsorbed ethanol, acetate, crotonaldehyde, and carbonate were identified as significant adsorbed species, and again no adsorbed CO was found. When the Rh modifier was added to the Co/ $\text{CeO}_2$ , the stability of ethoxide decreased, and linearly adsorbed CO appeared in the IR spectra between 373 and 673 K. The amount of acetate was lower, and in a narrow temperature range (423–573 K), acetaldehyde and traces of crotonaldehyde were observed. The promoting effect of Rh was mainly rationalized by an increased efficiency in C–C bond rupture on both Rh and metallic Co sites.

#### ■ ASSOCIATED CONTENT

##### Supporting Information

Additional conversion and selectivity data for the studied catalysts. This information is available free of charge via the Internet at <http://pubs.acs.org/>.

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##### Notes

The authors declare no competing financial interest.

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#### ■ REFERENCES

- (1) Brown, L. F. *Int. J. Hydrogen Energy* **2001**, *26*, 381–397.
- (2) Haryanto, A.; Fernando, S.; Murali, N.; Adhikari, S. *Energy Fuels* **2005**, *19*, 2098–2106.
- (3) Vaidya, P. D.; Rodrigues, A. E. *Chem. Eng. J.* **2006**, *117*, 39–49.
- (4) Ni, M.; Leung, Y. C.; Leung, M. K. H. *Int. J. Hydrogen Energy* **2007**, *32*, 3238–3247.
- (5) de la Piscina, P. R.; Homs, N. *Chem. Soc. Rev.* **2008**, *37*, 2459–2467.
- (6) Mattos, L. V.; Jacobs, B. H.; Noronha, F. B. *Chem. Rev.* **2012**, *112*, 4094–4123.
- (7) Yee, A.; Morrison, S. J.; Idriss, H. *Catal. Today* **2000**, *63*, 327–335.
- (8) Aupretre, F.; Descorme, C.; Duprez, D. *Catal. Commun.* **2002**, *3*, 263–267.
- (9) Song, H.; Ozkan, U. S. *J. Catal.* **2009**, *261*, 66–74.
- (10) Bayram, B.; Soykal, I. I.; Deak, D.; Miller, J. T.; Ozkan, U. S. *J. Catal.* **2011**, *284*, 77–89.
- (11) Song, H.; Mirkelamoglu, B.; Ozkan, U. S. *Appl. Catal., A* **2010**, *382*, 58–64.
- (12) Badlani, M.; Wachs, I. E. *Catal. Lett.* **2001**, *75*, 137–149.
- (13) Kulkarni, D.; Wachs, I. E. *Appl. Catal., A* **2002**, *237*, 121–137.
- (14) Erdőhelyi, A.; Raskó, J.; Kecskés, T.; Tóth, M.; Dömök, M. *Catal. Today* **2006**, *116*, 367–376.
- (15) Yee, A.; Morrison, S. J.; Idriss, H. *J. Catal.* **1999**, *186*, 279–295.
- (16) Lin, S. S.-Y.; Kim, D. H.; Ha, S. Y. *Catal. Lett.* **2008**, *122*, 295–301.
- (17) Llorca, J.; Homs, N.; Sales, J.; de la Piscina, P. R. *J. Catal.* **2002**, *209*, 306–317.
- (18) Llorca, J.; Homs, N.; de la Piscina, P. R. *J. Catal.* **2004**, *227*, 556–560.
- (19) Lin, S. S.-Y.; Kim, D. H.; Ha, S. Y. *Appl. Catal., A* **2009**, *355*, 69–77.
- (20) Ponec, V. *Stud. Surf. Sci. Catal.* **1991**, *64*, 117–157.
- (21) Guzzi, L.; Boskovic, G.; Kiss, E. *Catal. Rev.* **2010**, *52*, 133–203.
- (22) Johnson, B. G.; Bartholomew, C. H.; Goodman, D. W. *J. Catal.* **1991**, *128*, 231–247.
- (23) Ho, S.-W.; Houalla, M.; Hercules, D. M. *J. Phys. Chem.* **1990**, *94*, 6396–6399.
- (24) Khodakov, A. Y.; Griboval-Constant, A.; Bechara, R.; Zholobenko, V. L. *J. Catal.* **2002**, *206*, 230–241.
- (25) Kaspar, J.; Fornasiero, P.; Graziani, M. *Catal. Today* **1999**, *50*, 285–298.
- (26) *Catalysis by Ceria and Related Materials*; Trovarelli, A., Ed.; Imperial College Press: London, 2002.

- (27) Happel, M.; Mysliveček, J.; Johánek, V.; Dvorak, F.; Stetsovych, O.; Lykhach, Y.; Matolín, V.; Libuda, J. *J. Catal.* **2012**, *289*, 118–126.
- (28) da Silva, A. M.; de Souza, K. R.; Mattos, L. V.; Jacobs, G.; Davis, B. H.; Noronha, F. B. *Catal. Today* **2011**, *164*, 234–239.
- (29) Martono, E.; Vohs, J. V. *ACS Catal.* **2011**, *1*, 1414–1420.
- (30) Mavrikakis, M.; Barteau, M. J. *Mol. Catal. A* **1998**, *131*, 135–147.
- (31) Diagne, C.; Idriss, H.; Kiennemann, A. *Catal. Commun.* **2002**, *3*, 565–571.
- (32) Sheng, P.-Y.; Yee, A.; Bowmaker, G. A.; Idriss, H. *J. Catal.* **2002**, *208*, 393–403.
- (33) Raskó, J.; Hancz, A.; Erdőhelyi, A. *Appl. Catal., A* **2004**, *269*, 13–25.
- (34) Tóth, M.; Dömök, M.; Raskó, J.; Hancz, A.; Erdőhelyi, A. *Chem. Eng. Trans.* **2004**, *4*, 229–234.
- (35) Raskó, J.; Dömök, M.; Báán, K.; Erdőhelyi, A. *Appl. Catal., A* **2006**, *299*, 202–211.
- (36) Raskó, J.; Kiss, J. *Appl. Catal., A* **2005**, *287*, 252–260.
- (37) Weststrate, C. J.; Gericke, H. J.; Verhoeven, M. W. G. M.; Ciobica, I. M.; Saib, A. M.; Niemantsverdriet, J. W. J. *Phys. Chem. Lett.* **2011**, *1*, 1767–1770.
- (38) Mullins, D. R.; Senanayake, S. D.; Chen, T.-L. *J. Phys. Chem. C* **2010**, *114*, 17112–17119.
- (39) Martono, E.; Vohs, J. M. *J. Catal.* **2012**, *291*, 79–86.
- (40) Martono, E.; Hyman, M. P.; Vohs, J. M. *Phys. Chem. Chem. Phys.* **2011**, *13*, 9880–9886.
- (41) Hyman, M. P.; Vohs, J. M. *Surf. Sci.* **2011**, *605*, 383–389.
- (42) Batista, M. S.; Santos, R. K. S.; Assaf, E. M.; Assaf, J. M.; Ticianelli, E. A. *J. Power Sources* **2003**, *124*, 99–103.
- (43) Óvári, L.; Krick Calderon, S.; Lykhach, Y.; Libuda, J.; Erdőhelyi, A.; Papp, C.; Kiss, J.; Steinrück, H.-P. *J. Catal.* **2013**, *307*, 132–139.
- (44) Virgine, M.; Araque, M.; Roger, A.-C.; Vargas, J. C.; Kiennemann, A. *Catal. Today* **2008**, *138*, 21–27.
- (45) Profeti, L. P. R.; Ticianelli, E. A.; Assaf, E. M. *J. Power Sources* **2008**, *175*, 482–489.
- (46) Cook, K. M.; Poudyal, S.; Miller, J. T.; Bartholomew, C.; Hecker, W. C. *Appl. Catal., A* **2012**, *449*, 69–80.
- (47) Jongsomjit, B.; Panpranot, J.; Goodwin, J. G. *J. Catal.* **2001**, *204*, 98–109.
- (48) Jacobs, G.; Das, T. K.; Patterson, P. M.; Li, J.; Sanchez, L.; Davis, B. H. *Appl. Catal., A* **2003**, *247*, 335–343.
- (49) Schanke, D.; Vada, S.; Blekkan, E. A.; Hilmen, A. M.; Hoff, A.; Holmen, A. *J. Catal.* **1995**, *156*, 85–95.
- (50) Prieto, G.; Concepcion, P.; Murciano, R.; Martinez, A. *J. Catal.* **2013**, *302*, 37–48.
- (51) Pereira, E. B.; Homs, N.; Marti, S.; Fierro, J. L. G.; de la Piscina, P. R. *J. Catal.* **2008**, *257*, 206–214.
- (52) Prieto, G.; Martinez, A.; Concepcion, P.; Moreno-Tost, R. *J. Catal.* **2009**, *266*, 129–144.
- (53) Yao, H. C.; Yu Yao, Y. T. *J. Catal.* **1984**, *86*, 254–265.
- (54) Jacobs, G.; Ji, Y.; Davis, B. H.; Cronauer, D.; Kropf, A. J.; Marshall, C. L. *Appl. Catal., A* **2007**, *333*, 177–191.
- (55) Noronha, F. B.; Frydman, A.; Aranda, D. A. G.; Perez, C.; Soares, R. R.; Morawek, B.; Castner, D.; Campbell, C. T.; Frety, R.; Schmal, M. *Catal. Today* **1996**, *28*, 147–157.
- (56) Lin, S. S.-Y.; Kim, D. H.; Engelhard, M. H.; Ha, S. Y. *J. Catal.* **2010**, *273*, 229–235.
- (57) Chin, R. L.; Hercules, D. M. *J. Phys. Chem.* **1982**, *86*, 360–367.
- (58) Burroughs, P.; Hamnett, A.; Orchard, A. F.; Thornton, G. *J. Chem. Soc., Dalton Trans.* **1976**, *17*, 1686–1698.
- (59) Pfau, A.; Schierbaum, K. D. *Surf. Sci.* **1994**, *321*, 71–80.
- (60) Fujimori, A. *Phys. Rev. B* **1983**, *28*, 4489–4499.
- (61) Mullins, D. R.; Overbury, S. H.; Huntley, D. R. *Surf. Sci.* **1998**, *409*, 307–319.
- (62) Guil, J. M.; Homs, N.; Llorca, J.; de la Piscina, P. R. *J. Phys. Chem. B* **2005**, *109*, 10813.
- (63) Vargas, J. C.; Libs, S.; Roger, A. C.; Kiennemann, A. *Catal. Today* **2005**, *107*, 417.
- (64) Dömök, M.; Tóth, M.; Raskó, J.; Erdőhelyi, A. *Appl. Catal., B* **2007**, *69*, 262–272.
- (65) Raskó, J.; Kiss, J. *Catal. Lett.* **2005**, *101*, 71–77.
- (66) Óvári, L.; Solymosi, F. *Langmuir* **2002**, *18*, 8829–8835.
- (67) Di Cosimo, J. I.; Diez, V. K.; Xu, M.; Iglesia, E.; Apesteguia, C. *R. J. Catal.* **1998**, *178*, 499–510.
- (68) Christiansen, M. A.; Mpourmpakis, G.; Vlachos, D. G. *ACS Catal.* **2013**, *3*, 1965–1975.
- (69) Li, M.; Wu, Z.; Overbury, S. H. *J. Catal.* **2013**, *306*, 164–176.
- (70) Vayssilov, G. N.; Mihaylov, M.; St. Petkov, P.; Hadjivanov, K. I.; Neyman, K. M. *J. Phys. Chem. C* **2011**, *115*, 23435–23454.
- (71) Yee, A.; Morrison, S. J.; Idriss, H. *J. Catal.* **2000**, *191*, 30–45.
- (72) Idriss, H.; Diagne, C.; Hinderman, J. P.; Kiennemann, A.; Barteau, M. A. *J. Catal.* **1995**, *155*, 219–237.
- (73) Soykal, I. I.; Sohn, H.; Ozkan, U. S. *ACS Catal.* **2012**, *2*, 2335–2348.
- (74) Soykal, I. I.; Bayram, B.; Sohn, H.; Gawade, P.; Miller, J. T.; Ozkan, U. S. *Appl. Catal., A* **2012**, *449*, 47–58.
- (75) Singh, M.; Zhou, N.; Paul, D. K.; Klabunde, K. J. *J. Catal.* **2008**, *260*, 371–379.
- (76) Solymosi, F.; Pásztor, M. *J. Phys. Chem.* **1985**, *89*, 4789–4793.
- (77) Gazsi, A.; Koós, Á.; Bánsági, T.; Solymosi, F. *Catal. Today* **2011**, *160*, 70–80.
- (78) da Silva, A. M.; de Souza, K. R.; Jacobs, G.; Graham, U. M.; Davis, B. H.; Mattos, L. V.; Noronha, F. B. *Appl. Catal., B* **2011**, *102*, 94–109.
- (79) Zhang, J.; Zhong, Z.; Cao, X.-M.; Hu, P.; Sullivan, M. B.; Chen, L. *ACS Catal.* **2014**, *4*, 448–456.